Table IV yield —0.03767. The latter figure also gives the same potential (0.05468) for the systems:

$$Ag + AgCl, HCl\ (0.01665)\ \|\ HCl\ (0.001665), AgCl + Ag$$

and

$$Ag + AgCl, KCl\ (0.01665)\ \|\ KCl\ (0.001665), AgCl + Ag$$

as would be expected if hydrochloric acid and potassium chloride are equally dissociated, since all the electrodes are reversible to the chloride ion. These computations will be discussed further in a forthcoming article.

I am indebted to Prof. A. A. Noyes for valuable criticism in the preparation of this paper.

4. Summary.

Since the product of the transference number of the chloride ion and the equivalent conductance $n\Lambda$ is a constant, at a given concentration and temperature, for the chlorides of the alkali metals and of hydrogen, it has been assumed that the "activity" of the chloride ion in these solutions is also independent of the nature of the (univalent) positive ion. On this basis the activity coefficients of both ions of hydrochloric acid have been calculated for a series of concentrations, and the potential of the normal hydrogen electrode against the normal calomel electrode has been computed.

CAMBRIDGE, MASS.

NOTE.

Corrections.—In the paper entitled "Notes on Neutral Salt Catalysis" (THIS JOURNAL, 40, 1461 (1918)). On p. 1463, Equation 2 should read $H_2O_2 + IO' = H_2O + I' + O_2$, instead of $H_2O_2 + IO' = H_2O + IO'_2$; on p. 1477, in the middle of the page, for $E_l + E_t$ read $E_t - E_l$; in the last line, for $E_t + E_l$ read $E_t - E_l$; on p. 1478, in Table VII, all numbers in the column headed E_l should be positive instead of negative; the column headed $E_t + E_l$ should read $E_t - E_l$; and on p. 1479, in Table VIII, under the column headed $\frac{K \times 10^2}{a_1}$, read 14.14 for 13.14.

HERBERT S. HARNED.

JOHN HARRISON LABORATORY,
UNIVERSITY OF PENNSYLVANIA,
PHILADELPHIA.